Review Assignment

April 10, 2019 3:16 PM

REVIEW ASSIGNMENT: Part A + Part B Ionic & Covalent Naming and Formula writing Practice

This assignment is to be completed below in the space provided.

Part A) Ionic & Covalent Compound Formula Writing

First, identify whether these compounds are ionic or covalent. Then, use the correct formula writing rules to write the correct chemical formulas for each compound.

	Compound Name	Type of Compound: Ionic or Covalent
1)	copper (II) chlorite	I
2)	sodium hydroxide	I
3)	nitrogen dioxide	C
4)	cobalt (III) oxalate	I
5)	ammonium sulfide	I
6)	aluminum cyanide	I
7)	carbon disulfide	C
8)	tetraphosphorous pentoxide	C
9)	potassium permanganate	I
10)	manganese (III) chloride	I
	Compound Name	**Type of Compound: Ionic or Covalent**
11)	calcium bromate	I
12)	carbon monoxide	C
13)	potassium oxide	I
14)	antimony tribromide	I
15)	zinc phosphate	I
16)	copper (II) bicarbonate	I
17)	dinitrogen tetroxide	C
18)	manganese (IV) carbonate	I
19)	lead (IV) nitride	I
20)	pentacarbon decahydride	C

	Chemical Formula	
1)	copper (II) chlorite	$Cu(ClO)_2$
2)	sodium hydroxide	$NaOH$
3)	nitrogen dioxide	NO_2
4)	cobalt (III) oxalate	$Co_2(C_2O_4)_3$
5)	ammonium sulfide	$(NH_4)_2S$
6)	aluminum cyanide	$Al(CN)_3$
7)	carbon disulfide	CS_2
8)	tetraphosphorous pentoxide	P_4O_5
9)	potassium permanganate	$KMnO_4$
10)	manganese (III) chloride	$MnCl_3$
11)	calcium bromate	$Ca(BrO_3)_2$
12)	carbon monoxide	CO
13)	potassium oxide	K_2O
14)	antimony tribromide	$SbBr_3$
15)	zinc phosphate	$Zn_3(PO_4)_2$
16)	copper (II) bicarbonate	$Cu(HCO_3)_2$
17)	dinitrogen tetroxide	N_2O_4
18)	manganese (IV) carbonate	$Mn(CO_3)_2$
19)	lead (IV) nitride	Pb_3N_4
20)	pentacarbon decahydride	C_5H_{10}

Part B) Ionic & Covalent Compound Naming

First, identify whether these compounds are ionic or covalent. Then, use the correct naming rules to write the correct names for each compound.

	Chemical Formula	Type of Compound:	Compound Name

Part B) Ionic & Covalent Compound Naming

First, identify whether these compounds are ionic or covalent. Then, use the correct naming rules to write the correct names for each compound.

	Chemical Formula	Type of Compound: Ionic or Covalent		Compound Name	
21)	$CdBr_2$	I	21)	$CdBr_2$	cadmium bromide
22)	$Cr(Cr_2O_7)_3$	I	22)	$Cr(Cr_2O_7)_3$	chromium (VI) dichromate
23)	SBr_2	C	23)	SBr_2	sulfur dibromide
24)	$(NH_4)_2CrO_4$	I	24)	$(NH_4)_2CrO_4$	ammonium chromate
25)	CuO	I	25)	CuO	copper (II) oxide
26)	$Pt_3(PO_3)_4$	I	26)	$Pt_3(PO_3)_4$	platinum (IV) phosphite
27)	$Al(ClO_4)_3$	I	27)	$Al(ClO_4)_3$	aluminum perchlorate
28)	NH_3	C	28)	NH_3	ammonia
29)	$Ca(C_2H_3O_2)_2$	I	29)	$Ca(C_2H_3O_2)_2$	calcium acetate
30)	N_2O	C	30)	N_2O	dinitrogen monoxide
	Chemical Formula	Type of Compound: Ionic or Covalent			
31)	$V(SO_4)_2$	I	31)	$V(SO_4)_2$	vanadium (IV) sulfate
32)	Ag_2CO_3	I	32)	Ag_2CO_3	silver carbonate
33)	N_2S_3	C	33)	N_2S_3	dinitrogen trisulfide
34)	$FeSO_3$	I	34)	$FeSO_3$	iron (II) sulfite
35)	$Zn(NO_2)_2$	I	35)	$Zn(NO_2)_2$	zinc nitrite
36)	$C_6H_{12}O_6$	C	36)	$C_6H_{12}O_6$	glucose
37)	PCl_3	C	37)	$Ni(NO_3)_2$	nickel (II) nitrate
38)	$Mn(OH)_7$	I	38)	PCl_3	phosphorus trichloride
39)	$Ni(NO_3)_2$	I	39)	$Mn(OH)_7$	manganese (VII) hydroxide
40)	O_2	C	40)	O_2	oxygen

3 2